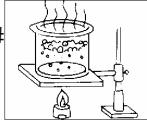

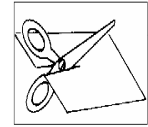
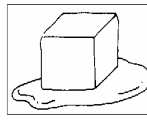


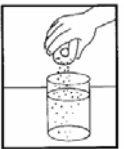
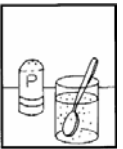
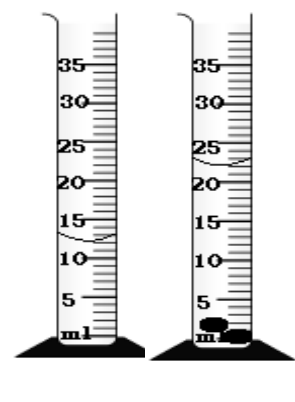


<p>1. All material that takes up space is called _____.</p> <p>\$ # matter % # liquid & # solid ' # gas (# volume</p>	<p>2. The amount of material in an object is called the object's _____.</p> <p>\$ # volume % # weight & # variety ' # mass (# density</p>
<p>3. An object's _____ is the pull of gravity on the object.</p> <p>\$ # density % # weight & # mass ' # volume (# solubility</p>	<p>4. A _____ has a definite volume but no definite shape.</p> <p>\$ # sublimation % # volume & # solid ' # liquid (# gas</p>
<p>5. A _____ does not have a definite volume or a definite shape.</p> <p># # gas % # liquid & # solid ' # condensation (# evaporation</p>	<p>6. A _____ has both a definite volume and shape.</p> <p>\$ # weight % # mass & # gas ' # liquid (# solid</p>
<p># #</p> <p>7. The amount of space a solid object takes up is its _____.</p> <p>\$ # weight % # mass & # volume ' # density (# reactivity</p>	<p># #</p> <p>8. # _____ can be used to identify a substance.</p> <p>\$ # volume % # density & # weight ' # mass (# color</p>
<p>9. _____ is when a liquid changes into a gas.</p> <p>\$ # condensation % # evaporation & # sublimation ' # melting (# freezing</p>	<p>10. # _____ is when a gas changes into a liquid.</p> <p>\$ # boiling % # melting & # condensation ' # evaporation (# sublimation</p>
<p>11. # _____ can be found without changing the substance into something else.</p> <p>\$ # creativity % # combustibility & # reactivity ' # chemical properties (# physical properties</p>	<p>12. To find the _____, divide the mass by the volume.</p> <p># #</p> <p>\$ # mass % # volume & # weight ' # density (# reactivity</p>
<p>13. The ability of a substance to go through a chemical change is called _____.</p> <p># #</p> <p>\$ # density % # reactivity & # combustibility ' # solubility (# conductivity</p>	<p>14. The ability of a substance to burn is called _____.</p> <p># #</p> <p>\$ # density % # reactivity & # combustibility ' # solubility (# conductivity</p>
<p>15. Which of the following is NOT a physical property?</p> <p>\$ # color % # density & # reactivity ' # solubility (# mass</p>	<p>16. When iron rusts, it no longer conducts electricity. This is because it _____.</p> <p>\$ # is combustible % # lost its luster & # is in water ' # changed physically (# changed chemically</p>

<p>17. The density of a object is a _____.</p> <p>\$ # physical property % # combustible property & # reactive property ' # chemical property (# soluble property</p>	<p>18. Weight is measured on a scale. Mass is measured on a ____.</p> <p>\$ # hygrometer % # barometer & # thermometer ' # scale (# balance</p>
<p>19. Changing the shape and amount of a substance does NOT change its _____.</p> <p>\$ # volume % # mass & # density ' # weight (# appearance</p>	<p>20. Which of the pictures shows a chemical change?</p> <p>\$ #  %  &  '  # (# None of the above</p>
<p>21. A person's _____ is different on the moon than on earth.</p> <p>\$ # mass % # weight & # volume ' # density (# size</p>	<p>22. A solution is a type of _____.</p> <p>\$ # mixture % # balance & # matter ' # combustibility (# chemical reaction</p>
<p>23. Which of the following does NOT signal a chemical change?</p> <p>\$ # color change % # production of light and heat & # production of gas ' # change in state (# change in combustibility</p>	<p>24. Separating red blood cells from plasma is an example of _____.</p> <p>\$ # using physical properties to separate a mixture % # using chemical properties to separate a mixture & # solubility ' # reactivity (# combustibility</p>
<p>25. What do the pictures show about the solubility of these substances?</p> <p>#  #  #  # </p>	<p>\$ # Pepper is more soluble than sugar. % # If stirred, the pepper would have the same solubility as sugar. & # Sugar and pepper have the same solubility. ' # Pepper is soluble in water, but sugar is not. (# Sugar is soluble in water, but pepper is not.</p>
<p>26. If the density of 3 grams of diamond is 3.51 g/cm³, what is the density of 6 grams of diamond?</p> <p>\$ # 21.06 g/cm³ % # 10.53 g/cm³ & # 7.02 g/cm³ ' # 3.51 g/cm³ (# 2 g/cm³</p>	<p>27. What is the volume of 1 marble?</p> <p>\$ # 22 ml % # 10 ml & # 5 ml ' # 2 ml (# 1 ml</p> 
<p>28. If the volume of an object is 4cm³, and its mass is 8g, what is its density?</p> <p>\$ # 32 g/cm³ % # 2 g/cm³ & # 4 g/cm³ ' # 12 g/cm³ (# 0.5 g/cm³</p>	<p>29. The change from a solid to a gas is _____.</p> <p>\$ # condensation % # evaporation & # sublimation ' # boiling (# melting</p>
<p>30. The melting point of iron is 1535°C. The freezing point of iron _____.</p> <p>\$ # is greater than 1535°C % # is less than 1535°C & # is 1535°C ' # cannot be determined from the given information. #</p>	<p>31. Oxygen can be compressed and cooled into a liquid at a very low temperature. So, oxygen normally exists as a ____.</p> <p>\$ # gas % # solid & # liquid ' # all of the above (# none of the above</p>

32. How does the volume of liquid water change when it turns to water vapor? **EXPLAIN!**

33. Can water be used to separate salt from sugar? **WHY OR WHY NOT?**

34. Describe what would happen to a cold empty glass if it was suddenly placed in a hot, humid room. **EXPLAIN YOUR ANSWER!**

35. Suppose you had 5g of paper. You burned it and found that its new mass was 3g. **Infer** where the 2g of mass went.

36. How do you measure density?
